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Chapter 5 Electrons In Atoms

138 Chapter 5 •

Electrons in Atoms

Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and

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Electrons In Atoms; Answers To Worksheet

frequency are inversely related; in other words, as one quantity increases, the other decreases.

Chapter 5: Electrons in Atoms

Replaces Bohr Model: electrons, instead of traveling in defined orbits or hard, spherical "shells," as Bohr proposed, travel in diffuse clouds around the nucleus.

Electron Configuration

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(5.3) Each orbital within a sublevel can only hold two (2) electrons. The electrons within an orbital must have opposite spins.

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Chapter 5 - Electrons
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Models of the Atom
The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom.

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Chapter 5: Electrons in Atoms. continued their quest to understand atomic structure and the arrangement of

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electrons within atoms. Rutherford proposed that all of an atom's positive charge and virtually all of its mass are concentrated in a nucleus that is surrounded by fast-moving electrons.

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Atoms of chlorine, a yellow-green gas at room temperature, react readily with atoms of many other

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elements. Figure 5-1 shows chlorine atoms reacting with steel wool. The interaction of highly reactive chlorine atoms with the large surface area provided by the steel results in a vigorous reaction.

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Chapter 5.1 to 5.3

Electrons In Atoms

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Terms in this set (...)

amplitude. the height
of a wave from the
origin to a crest, or
from the origin to a
trough. atomic

emission spectrum. the
set of frequencies of the
electromagnetic wave
emitted by atoms of
the element. atomic

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orbital. electrons
probable location.
aufbau principle.

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Electrons in Atoms

Electron Configurations

for Elements in Period

Three Table 5-4 Figure

5-19. This sublevel

diagram shows the

order in which the

orbitals are usually

filled. The proper

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sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d.

Chapter 5: Electrons in Atoms

Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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frequency are inversely related; in other words, as one quantity increases, the other decreases.

Chemistry Chapter 5 Electrons In Atoms Test Answers

Key Concepts Chapter 5 electrons in atoms answer key study guide. Atoms are made of extremely tiny particles called protons, neutrons, and electrons. Protons and

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neutrons are in the center of the atom, making up the nucleus
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Chapter 5 Electrons In Atoms Answer Key Study Guide

This video describes light as a particle and wave. It also describes matter and quantum of energy.

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Schiller. STUDY.

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Electrons in atoms.
Terms in this set (18)
energy levels. fixed
energies an electron
can have (circles
around nucleus)
quantum. the amount
of energy needed to
move an electron from
one energy level to ...

**Chemistry Chapter 5
Electrons in atoms-**

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...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:
Aufbau principle - electrons enter the lowest energy first.;
This causes difficulties because of the overlap of orbitals of different energies - follow the diagram!

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Chemistry: Matter and
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Supplemental Problems

1. Orange light has a
frequency of 4.8×10^{14}
 s^{-1} . What is the energy
of one quantum of
orange light?

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